

Chapter 17 Thermochemistry Test A Answers

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Thermochemistry - Chapter 17. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. Shelby_Southerland9. Terms in this set (20) endothermic. a chemical or physical process that takes in heat. exothermic. a chemical or physical process that releases heat. enthalpy. the heat content of a system.

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Chapter 17 "Thermochemistry" Tools. Copy this to my account; E-mail to a friend; Find other activities; Start over; Help; Use these activities to learn the vocabulary and major concepts presented in this chapter. A B; calorie: the quantity of heat that raises the temperature of 1 g of pure water by 1°C:

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Chapter 17 Thermochemistry. Matching (1 point each) Match each item with the correct statement below. a. calorimeter d. enthalpy b. heat capacity

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e. specific heat c. joule ____ 1. SI unit of energy ____ 2. quantity of heat needed to change the temperature of 1 g of a substance by 1 C ____

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Chapter 17 Thermochemistry Practice Problems

Chapter 17 Thermochemistry 187 10. Complete the enthalpy diagram for the combustion of natural gas. Use the thermochemical equation in the first paragraph on page 517 as a guide. SECTION 17.3 HEAT IN CHANGES OF STATE (pages 520-526) This section explains heat transfers that occur during melting, freezing, boiling, and condensing.

SECTION 17.1 THE FLOW OF ENERGY HEAT AND WORK (pages 505-510)

Thermochemistry Test Preview Matching Match each item with the correct statement below. a. calorimeter d. enthalpy b. calorie e. specific heat c. joule f. heat capacity ____ 1. quantity of heat needed to raise the temperature of 1 g of water by 1 C 2. ____ SI unit of energy 3. ____ quantity of heat needed to change the temperature of 1 g of a ...

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Review Unit 4 Test IPOD #32 - Chapter 13 concepts Review - Phase Changes Slide 24: Describe terms for phase changes Worksheet: Chapter 17 - Thermochemistry Notes Chapter 17 Notes, Slides 1-3: What is thermochemistry? Chapter 17 Notes, Slides 4-11: Heat, energy, calorie/joule conversions Chapter 17 Notes, Slides 12-13: Endo vs. Exo

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162 CHAPTER 6: THERMOCHEMISTRY To convert the answer to joules, we write: $101.3 \text{ J} \cdot 0.18 \text{ L atm} \cdot 1 \text{ L atm}^{-1} = -18 \text{ J}$. An expansion implies an increase in volume, therefore w must be -325 J (see the defining equation for pressure-volume work.) If the system absorbs heat, q must be $+127 \text{ J}$. The change in energy (internal

CHAPTER 6 THERMOCHEMISTRY

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