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3.zinc + hydrochloric acid (zinc chloride + hydrogen. 4.iron + copper(II) sulfate (iron (II) sulfate + copper. 5.copper + sulfuric acid (copper (II) sulfate + hydrogen. 6.potassium dichromate + sulfur + water (sulfur dioxide + potassium hydroxide + chromium (III) oxide.

BALANCING CHEMICAL EQUATIONS - PRACTICE SHEET # 2

Unformatted text preview: CHEMISTRY - CHAPTER 3: EQUATIONS BALANCING EQUATIONS: FORMULAS GIVEN Practice Sheet #1 Balance the following equations: 1. $\text{Al} + \text{N}_2 \rightarrow$ 2. $\text{Fe} +$ 3. $\text{CaCO}_3 \rightarrow$ 4. $\text{NH}_4\text{NO}_3 \rightarrow \text{N}_2\text{O} + \text{H}_2\text{O}$ 5. KI 6. $\text{Pb}(\text{NO}_3)_2 + \text{HCl} \rightarrow$ 7. $\text{BaO}_2 \rightarrow$ 8. $\text{Al} +$

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H₂SO₄ → 9. CH₄ + Cl₂ → 10. MgCl₂ +
NaOH → Mg(OH)₂ 11. AgNO₃ + CuCl₂ →
AgCl 12. ZnS 13. Na 14.

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CHAPTER 3: EQUATIONS BALANCING
EQUATIONS: FORMULAS GIVEN Practice
Sheet #1 Balance the following
equations: 1. Al + N₂ → 2. Fe + 3.
CaCO₃ → 4. NH₄NO₃ → N₂O + H₂O 5. KI
6. Pb(NO₃)₂ + HCl → 7. BaO₂ → 8. Al +
H₂SO₄ → 9. CH₄ + Cl₂ → 10. MgCl₂ +
NaOH → Mg(OH)₂ 11. AgNO₃ + CuCl₂ →
AgCl 12. ZnS ...

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Worksheet 1

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Worksheet #3: Decomposition Reactions
In decomposition reactions, one compound will break down into two or more parts. 1. barium carbonate \rightarrow barium oxide + carbon dioxide $\text{BaCO}_3 \rightarrow \text{BaO} + \text{CO}_2$ 2. magnesium carbonate \rightarrow magnesium oxide + carbon dioxide $\text{MgCO}_3 \rightarrow \text{MgO} + \text{CO}_2$ 3.

Worksheet #3: Decomposition Reactions

Stoichiometry, part 3; Chapter 5 (Introduction to Solutions and Aqueous Reactions) Solution concentration; Dilutions; Making solutions; Molarity;

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Titration; Molecular, Full, and Net-ionic equations; Chapter 6 (Gases) Gas laws; Gas stoichiometry; Chapter 7 (Thermochemistry) Calorimetry worksheet; Calorimetry exercises and answers

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Equation 11.3.2 is identical to Equation 11.3.1 except for the addition of abbreviations in parentheses to indicate the physical state of each species. The abbreviations are (s) for solid, (l) for

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liquid, (g) for gas, and (aq) for an aqueous solution, a solution of the substance in water.. Consistent with the law of conservation of mass, the numbers of each type of atom are the same on both ...

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